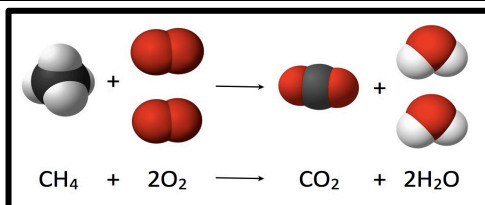
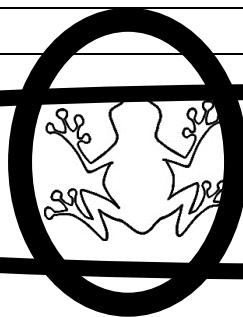


Name

Period



Methane Gas + Oxygen → Carbon Dioxide + Water



OBJECTIVE:

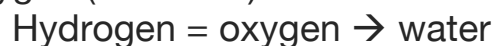
To MODEL evidence for the Law of Conservation of Matter & Mass during Chemical Reactions.

- ___1) Use chemical equations illustrated with diagrams of atoms and molecules. Use a different color for each element.
- ___2) Include a KEY, a COUNT of atoms and the total MASS for each reactant and product.

Background:

- ___1) Atoms of reactants are regrouped into different molecules, resulting in products and are written as a chemical equation.
- ___2) The total number of each type of atom, and the total mass, stays the same (Laws of Conservation of Matter and of Mass).
- ___3) The new substances (products) have different properties than the properties of the original substances (reactants).

1. Hydrogen burning in oxygen (oxidation):



2. Photosynthesis: Plants change the energy of the sun to glucose (sugar or food):



3. Respiration: Animals use glucose (sugar or food) and oxygen (oxidation) for energy:



4. Nitrogen and Hydrogen combine to form Ammonia.



5. Sodium burning in chlorine forms sodium chloride.